

Applied Adsorption Kinetics Model for Removal of Hazards from Aqueous Solution: More Informational Parameters for Industrial Design

Song Lou*

Department of Biological Systems Engineering, Virginia Polytechnic Institute and State University, Blacksburg, VA 24061, United States

***Corresponding Author:** Song Lou, Department of Biological Systems Engineering, Virginia Polytechnic Institute and State University, Blacksburg, VA 24061, United States, Tel: +15402317601; E-mail: songl@vt.edu

Received: 18 September 2017; **Accepted:** 22 September 2017; **Published:** 06 October 2017

Abstract

Adsorption is one of the widely applied techniques for removal of hazards from aqueous solution. Its kinetics studies are of great significance to evaluate the adsorption performance and reveal the underlying mechanisms. There are several mathematic models to describe the adsorption process and diffusion mode. In recently years, more and more adsorption kinetics models with more informational parameters are investigated, which could give guidance for industrial conversion through supplying paramount importance parameters. Here, we reviewed several widely used adsorption kinetics models and paid more attention to their potential values of industrial application guiding. We believe that the review is of certain significance and improvement of applied adsorption kinetics modeling.

Keywords: Adsorption kinetics models; Industrial conversion; Wastewater treatment; Kinetics parameters

1. Removal of Hazards from Aqueous Solution

Recent advances in environment science and engineering suggest that many of the current problems involving water quality could be greatly diminished using porous adsorption resin [1]. It has recently been demonstrated that porous-structured materials and particle exhibit good adsorption efficiency especially due to their high surface area and greater amount of active sites for interaction with toxic metallic species [2-8]. However, there is less attention on the potential practice-guiding impact of related adsorption model. A variety of adsorption kinetic and induction force

models, like pseudo first order, Langmuir and Freundlich models, which are widely used in the treatment study of wastewater, are used to try to investigate the mechanism of adsorption process [9-15]. In recent years, the study of adsorption models, which could indicate the applied parameters, such as adsorption equilibrium time and optimal adsorbent pore size, is being researched extensively [16-18]. These models are expected to be applied prosperously since the supplied parameters are very vital to optimize the design of the industrial adsorption system.

As scientists begin to identify the adsorption mechanisms that contribute to separation efficiency, more informed and personalized removal strategy of hazard from aqueous solution can be developed. This increased insight has led to the directed design of adsorption kinetics models that better exploit the physical and chemical properties of adsorbent. For instance, in industrial scale objects, the adsorbent and adsorbate are in contact with each other for a short time period, therefore, predicating the equilibrium time and the rate of adsorption is of paramount importance [19-22]. Meanwhile, differences in particle size and the pore nature of the adsorbent plays a very important role in the efficiency of the adsorption process [23-26]. The better description of adsorbent size contribution could supply a better guidance for the adsorbent size selection in big-scale application of water treatment.

Many hypotheses have been considered to describe the adsorption kinetic and various theoretical and empirical models have been presented for modeling of adsorption kinetics. The classical adsorption kinetics model, Langmuir kinetics model (equation 1): assumes the surface of adsorbate is energetically homogeneous [27-28].

$$\theta_A = \frac{V}{V_m} = \frac{K_{eq}^A p_A}{1 + K_{eq}^A p_A} \quad (1)$$

(p_A is adsorbate's partial pressure; V is the volume of adsorbate; θ_A is the fractional occupancy of the adsorption sites; V_m is the volume of the monolayer; K_{eq} is the associated equilibrium constant). By contrast, the Langmuir-Freundlich model (equation 2) is another equation, which can't be solved analytically and considers the effect of surface heterogeneity [29-31]. However, the rate constant of this model can be obtained by an approximation method.

$$\theta_A = \alpha_F p^{C_F} \quad (2)$$

(θ_A is the fractional occupancy of the adsorption sites; α_F and C_F are fitting parameters)

The other two important empirical rate models: pseudo first order (equation 3) and pseudo second order (equation 4): could supply a better validity under certain conditions [32-34].

$$\frac{dq}{dt} = k_1(q_e - q) \quad (3)$$

(k_1 is the pseudo first order rate coefficient; q_e and q are the amount of the adsorbed species per unit mass of adsorbent at the equilibrium state and any time of adsorption)

$$\frac{dq}{dt} = k_2(q_e - q)^2 \quad (4)$$

(k_2 is the pseudo second order rate coefficient)

All these classical adsorption kinetics models are used generally in the study of wastewater treatment to investigate the adsorption mechanism [21, 35-39]. However beside the traditional adsorption theory, they cannot give any specific guidance parameter for the industry-scale application. As more and more lab research of wastewater began to be transformed via industrial conversion, more informative process parameters needs to be developed from kinetics model to guide industrial production. In 2017, Azizian et al. reported that the shape of the adsorption kinetic curve could be described by the mathematical model of hyperbolic tangent properly (equation 5) [16] Followed by the traditional model deduction process, the y and x were defined as $y = \frac{q}{q_e}$ and $x=kt$. When the equilibrium state is set as $q/q_e \geq 0.995$, the adsorption kinetics model could be described as $\frac{q}{q_e} = \tanh(\pi \frac{t}{t_e})$. In addition to consider the different degrees of curvature of kinetics curves, which are derived from the adsorbent surface heterogeneity, different functional groups on the ad-species, possibility of binding to the surface with different geometries, and effect of diffusion on the adsorption overall rate, the n power law was introduced for the curvature degree:

$$\frac{q}{q_e} = [\tanh(\pi \frac{t}{t_e})]^n \quad (5)$$

(t_e is the required time for adsorption to reach the equilibrium state; π is the setting of $q/q_e \geq 0.995$; n is power constant)

This hyperbolic tangent model was evaluated through three different data series on two different testing conditions: adsorption same concentration of adsorbate onto the different amount of adsorbent; different concentration of adsorbate onto the same concentration of adsorbent. For the different amount of adsorbent, the values of the equilibrium time derived from hyperbolic tangent model decreased by increasing the adsorbent dosage. For the different amount of adsorbate, the equilibrium time, t_e , increased by increasing the adsorbate concentration in solution.

The predicted equilibrium times of adsorption by the hyperbolic tangent kinetic model are very similar to the experimental ones. This model is simple and can be used easily for modeling of adsorption kinetics data in industrial scales to give the equilibrium time of adsorption. This predicted equilibrium time is much useful for controlling the industrial wastewater treatment process, which could benefit the industrial design and production cost. It is a perfect example of how basic science of model and engineering go hand in hand. This work investigated the model with the restriction setting of equilibrium state as $q/q_e > 0.995$. Actually, this model cannot only be used under this scale, but

also any designed equilibrium degree. So it could accurately predict the equilibrium time in any setting scale of q/q_e according to the experiment needed. By leveraging our ever-growing knowledge of wastewater adsorption and its kinetics model, it is possible to engineer a progressive hyperbolic tangent kinetics models to give more industrial-guiding parameters under more flexible equilibrium definition.

Recently, rates of adsorption have been observed to exhibit biphasic behavior: a rapid adsorption in the order of minutes followed by a slower phase with a timescale in the order of hours [40-41]. The latter phase, which may result in a significant fraction being sequestered, [42-44] is frequently rate limiting for abstraction and separation [45-47]. Mechanism proposed as potentially responsible for the commonly observed slow adsorption of hazards on resins includes intraorganic matter diffusion [48, 49] and hindered pore diffusion [50-52]. Diffusion through hazardous materials is analogous to diffusion in polymers; i.e, diffusing molecules must penetrate and migrate through a polymeric matrix [53]. Pore diffusion can be either sterically hindered or retarded by sorption to organic phases associated with pore wells. Adsorption on high-energy sites may also be an important rate-limiting step. In condensed polymeric organic matter, active compounds may adsorb strongly in molecule-sized voids or “holes” [54] while during pore diffusion they are subject to high-energy sorption in molecule-sized pores [55-56]. Now, adsorption has been modeled both as a diffusive and as a first-order rate process. Several different rate models have been used to describe the adsorption of active compounds onto adsorbent.

Different groups have carried out adsorption kinetics studies. Norris et al. presented the first experimental evidence to produce a three-parameter model (equation 6) to describe adsorption rates for both a slowly adsorbing fraction and a rapidly desorbing fraction [57-58]. The model’s parameters are useful for distinguishing the rapidly and slowly adsorbing compartments. An advantage of this model is that its parameters are useful for distinguishing between rapidly and slowly adsorbing fractions.

$$\frac{q(t)}{q_0} = \phi_s \exp(-k_s t) + (1 - \phi_s) \exp(-k_r t) \quad (6)$$

(ϕ_s is fast adsorption rate coefficient; k_s is fast adsorption rate; k_r is slow adsorption rate; t is adsorption time, $q(t)$ is adsorption amount (mg/g) at the time of t ; q_0 is initial adsorption amount(mg/g).

Furthermore, model results are intuitive given the biphasic nature of the adsorption data. Noort et al. stated a technique using Tenax TA beads as a “sink” for the solute was employed to measure the kinetics of adsorption and desorption of chlorobenzenes, polychlorinated biphenyls from laboratory-contaminated sediment [59]. They suggested a three-compartment model (equation 7) in which the slow fraction is divided into two compartments, slowly and very slowly adsorbing. An advantage of this model is that it theoretically describes a continuous distribution of sorptive compartments with only two parameters. However, convergence on a five-parameter solution was laborious and uncertain, with final fitted parameters strongly depending on initial guesses; researchers rarely recommend the triphasic model over others.

$$\frac{q(t)}{q_0} = \phi_r \exp(-k_r t) + \phi_s \exp(-k_s t) + (1 - \phi_r - \phi_s) \exp(-k_{vs} t) \quad (7)$$

(ϕ_s is fast adsorption rate coefficient; k_s is fast adsorption rate; k_r is slow adsorption rate; t is adsorption time, $q(t)$ is adsorption amount (mg/g) at the time of t ; q_0 is initial adsorption amount (mg/g))

Shuler studied a new two-parameter model (equation 8) to express the kinetics of adsorption based on the gamma function, which assumes the adsorption and desorption from heterogeneous sites occurs over a continuum of energies and rates [60]. This model theoretically describes a continuous distribution of adsorptive compartments with only two parameters. An advantage of this model is that it theoretically describes a continuous distribution of sorptive compartments with only two parameters. However, the gamma model underpredicted adsorption at late stages as it “flattened out” too much, indicating that the model does not mechanistically describe the observed adsorption process; the two calculated parameters do not have obvious physical significance and do not logically correlate to observed trends in slow adsorption rates and resistant fractions.

$$D_{app} = \frac{D_{aq}\epsilon_i}{(\epsilon_i + K_d \rho)\tau_f} \quad (8)$$

(D_{aq} is the diffusion coefficient; ρ_g is the density of the adsorbate; ϵ_i is the intraparticle porosity; τ_f is the tortuosity factor)

Roberts and Gschwend used a one-parameter spherical pore-diffusion model (equation 9) to describe the experimental data of adsorption, which only has one fitting parameter, and it can theoretically estimate adsorption rates a priori based on adsorbate and adsorbent physical characteristics [61]. Advantages of this model are that it has only one fitting parameter, and it can theoretically estimate adsorption rates a priori based on sorbate and sorbent physical characteristics. However, the simple spherical pore-diffusion model often does not fit entire adsorption rate profiles, and it essentially has no rapidly adsorbing fraction.

$$\frac{\partial c_r}{\partial r} = D_a \frac{\partial}{\partial r} \left(r^2 \frac{\partial c_r}{\partial r} \right) \quad (9)$$

(C_r is the immobile aqueous-phase concentration, r is the radial coordinate; D_a is the apparent diffusion coefficient)

Although there are several reports about the adsorption kinetics of resins, all of them attribute the adsorption to a simple process without considering the effect of sphere size. Corroll and Berens suggested a three-parameter, two-compartment model (equation 10) accounting for the influence of sphere diameter, which assumed biphasic polymer diffusion processes [62]. They studied rates of adsorption of solutes to poly-(vinyl chloride) polymer powders of non-uniform particle size and found that a simple modification of the Fickian model satisfactorily accounts for the effect of particle non-uniformity upon sorption kinetics for conditions under which transport is diffusion-controlled.

This monophasic model always gives a good fit of data in each case, however the adsorption is not considered as rapid, slow and even slower procedures, but only a global process instead.

$$\frac{q(t)}{q_0} = \frac{6}{\pi^2} \sum_{n=1}^{\infty} \frac{1}{n^2} [\phi_r \exp\left(\frac{-4n^2\pi^2 D_r t}{a_r^2}\right) + (1 + \phi_r) \exp\left(\frac{-4n^2\pi^2 D_s t}{a_s^2}\right)] \quad (10)$$

(ϕ_r is fast adsorption rate coefficient; a_r and a_s are the diameters of the spheres; D_r and D_s are diffusion coefficient)

My study was performed in order to determine the adsorption profiles of adsorbate on resins functionalized with chloromethyl and amino groups [63-64]. A new adsorption isotherm model, the multi-layer inductive effect model (equation 11): in which the inductive effect passed to a higher layer, and a new adsorption kinetics model (equation 12): the multi-compartment kinetics model, in which the adsorption process contained more compartments, were created according to the multiparameter theory and Karickhoff's theory by investigating the regression of the experimental results [65].

$$\frac{Q_e}{Q_m} = \frac{K_i C_e \{ (1 - K_i C_e) \sum_{i=3}^i [K_1 C_e + 2K_1 K_2 C_e^2 + \dots + (i-3)K_1 K_2 \dots K_{i-2} C_e^{i-1}] + (i-2)K_1 K_2 \dots K_{i-2} C_e^{i-1} (1 + K_i C_e - K_{i+1} C_e) \} + K_1 K_2 \dots K_i C_e^{i-1}}{(1 - K_i C_e) [(1 + K_{i-1} C_e) \sum_{i=3}^i (1 + K_i C_e - K_{i+1} C_e^2) + K_1 K_2 \dots K_{i-2} C_e^{i-2} (1 + K_{i-1} C_e - K_i C_e)]} \quad (11)$$

(K_1, K_2, \dots, K_i are the equilibrium adsorption constants for the first layer, second layer, ..., i th layer, respectively; $\theta_1, \theta_2, \dots, \theta_i$ are the fractions of surface coverage at the first layer, second layer, ..., i th layer, respectively; C_e is the equilibrium adsorbate concentration; Q_m is the surface site density)

$$\frac{M_t}{M_0} = F_1 e^{-k_1/(k_1'+k_1)t} + F_2 e^{-k_1 k_2/(k_1'+k_2)t} + F_3 e^{-k_1 k_2 k_3/(k_1'+k_2)(k_2'+k_3)t} + \dots + F_i e^{-\frac{k_1 k_2 k_3 k_i}{(k_1'+k_2)(k_2'+k_3) \dots (k_{i-1}'+k_i)} t} \quad (12)$$

(F_1, F_2, \dots, F_i are different compartments fractions; k_1, k_2, \dots, k_i are the rate constants of every adsorption level; k_1', k_2', \dots, k_i' are the rate constants of every desorption level, respectively; M_0 is mass present at $t=0$)

These new models explained the adsorption process better than the other kinetic models, and provided a good fit of adsorption data, which is not surprising in that they had at least two more fitting parameters than any of the other models. In order to guide the industry application, a sphere size parameter was considered to be included to explain adsorption kinetics more explicitly. Previous studies have demonstrated that due to the repulsive or attractive action between adsorbate and adsorbent, adsorption of multiple components can be rationally expected to be either competitive or synergistic. In some studies, [41, 45, 56, 66] weeks or months of incubation for the adsorption systems were tested in order to make sure the sorption equilibrium was achieved. In fact, some previous studies have shown that aging a long time has negligible additional effect on adsorption-resistant fractions or adsorption rates. The equilibration time is based on several factors, such as driving force, sphere diameter, and the concentration of the solution. The effects of driving force, solution concentration and sphere size on the adsorption process have been discussed in our previous work [67].

Comparison of experimental data with Fick's Law requires a solution of the diffusion equation appropriate for the sample geometry and experimental boundary conditions. The simplest geometry for powder is a collection of spherical particles of uniform size. This condition can, in fact, be closely approached in specific powder prepared by emulsion polymerization. Consequently, to approximate adsorption of adsorbate by synthetic resins, we made a model for the adsorption on nonuniform resins in the case of uniform initial concentration throughout the sphere, a constant concentration at the surface, and constant D_r . According to the theory of the first-order, two-component four-parameter model, a new model [68] incorporating sphere size can be described by the following equation 13:

$$\frac{M_t}{M_0} = F_1 e^{-4 \times 2^2 \pi^2 D_{r_1} t / r_1^2} + F_2 e^{-4 \times 2^2 \pi^2 D_{r_2} t / (r_1 + r_2)^2} \quad (13)$$

(M_t is the solid-phase sorbate concentration at a given time, M_0 is the initial solid-phase adsorbate concentration, F_1 and F_2 are different compartments' fractions, D_r is the diffusion coefficient, and r_1 and r_2 are the diameters of the spheres, on which the different compartments of adsorption process are mainly carried out)

It should be noted, however, that this first-order, two-compartment, four-parameter model does not necessarily reflect the reasonable procedures of adsorption. This model contains just two steps, and does not take into account the differences between fast adsorption, slow adsorption, and even slower adsorption [69]. Accordingly, a modified model with sphere size was created according to Karickhoff's theory and the multiparameter adsorption kinetics model. A sphere-size model in which the adsorption process contains more compartments was proposed as equation 14.

$$\frac{M_t}{M_0} = F_1 e^{-4 \times n^2 \pi^2 D_{r_1} t / r_1^2} + F_2 e^{-4 \times n^2 \pi^2 D_{r_2} t / (r_1 + r_2)^2} + F_3 e^{-4 \times n^2 \pi^2 D_{r_3} t / (r_1 + r_2 + r_3)^2} + \dots + F_n e^{-4 \times \frac{n^2 \pi^2 D_{r_n} t}{(r_1 + r_2 + r_3 + \dots + r_n)^2}} \quad (14)$$

The terminated layer of the adsorption was determined by the multilayer polarity inductive effect model (equation 11).

Through studying the adsorption processes of same adsorbate on different resins and comparing the fitting results of different models, the values of R^2 from the sphere-size kinetics were identical with our previous multiparameter kinetics model. Meanwhile, this model supplied very useful parameters, which reflects the effect of sphere-size distribution on adsorption kinetics process. This sphere-size model was created to present the first experimental evidence that sphere size affects the adsorption process. The new sphere-size model has more detailed parameters to more accurately describe the adsorption process, and will assist in the development of an efficient, industrial wastewater purification process. However, this model still lacks several important parameters, which could guide the industrial-scale production better, such as equilibrium time.

Overall, the progress toward adsorption kinetics models with more detailed parameter is encouraging, and this is hope that further development of new adsorption theory will help to bring forth better industrial outcomes.

References

1. Tratnyek PG, Johnson RL. Nanotechnologies for environmental cleanup. *Nano today* 1 (2006): 44-48.
2. Diallo MS, Savage N. Nanoparticles and water quality. *Journal of Nanoparticle Research* 7 (2005): 325-330.
3. Mazzaro R, Boni A, Valenti G, et al. Uniform Functionalization of High-Quality Graphene with Platinum Nanoparticles for Electrocatalytic Water Reduction. *Chemistryopen* 4 (2015): 268-273.
4. Hwang S, Martinez D, Perez P, et al. Effect of surfactant-coated iron oxide nanoparticles on the effluent water quality from a simulated sequencing batch reactor treating domestic wastewater. *Environ. Pollut* 159 (2011): 3411-3415.
5. Jones LA, Ogden KL. Silica and titania nanoparticles impact on water quality: Experiments involving *Ralstonia pickettii* in nutrient-rich and poor media. *Environ Prog Sustain* 32 (2013): 279-284.
6. Zaluzhna O, Li, Y, Allison, TC, et al. Inverse-Micelle-Encapsulated Water-Enabled Bond Breaking of Dialkyl Diselenide/Disulfide: A Critical Step for Synthesizing High-Quality Gold Nanoparticles. *JACS* 134 (2012): 17991-17996.
7. Son J, Vavra J, Forbes VE. Effects of water quality parameters on agglomeration and dissolution of copper oxide nanoparticles (CuO-NPs) using a central composite circumscribed design. *Sci. Total Environ* 521 (2015): 183-190.
8. Bielefeldt AR, Stewart MW, Mansfield E, et al. Effects of chlorine and other water quality parameters on the release of silver nanoparticles from a ceramic surface. *Water Res* 47 (2013): 4032-4039.
9. Jewell KS, Castronovo S, Wick A, et al. New insights into the transformation of trimethoprim during biological wastewater treatment. *Water Res* 88 (2016): 550-557.
10. de Souza SMDGU, Bonilla KAS, de Souza AAU, et al. Removal of COD and color from hydrolyzed textile azo dye by combined ozonation and biological treatment. *J. Hazard. Mater* 179 (2010): 35-42.
11. Kluczka J, Korolewicz T, Zolotajkin M, et al. Boron removal from water and wastewater using new polystyrene-based resin grafted with glycidol. *Water Resour Ind* 11 (2015): 46-57.
12. Lakshmanan R, Sanchez-Dominguez M, Matutes-Aquino JA, et al. Removal of Total Organic Carbon from Sewage Wastewater Using Poly(ethylenimine)-Functionalized Magnetic Nanoparticles. *Langmuir* 30 (2014): 1036-1044.
13. Efremenko, I, Sheintuch, M, Predicting solute adsorption on activated carbon: Phenol. *Langmuir* 22 (2006): 3614-3621.
14. Singh KP, Mohan D, Sinha S, et al. Color removal from wastewater using low-cost activated carbon derived from agricultural waste material. *Industrial & Engineering Chemistry Research* 42 (2003): 1965-1976.
15. Amin MT, Alazba AA, Shafiq M. Adsorptive Removal of Reactive Black 5 from Wastewater Using Bentonite Clay: Isotherms, Kinetics and Thermodynamics. *Sustainability-Basel* 7 (2015): 15302-15318.
16. Eris S, Azizian S. Analysis of adsorption kinetics at solid/solution interface using a hyperbolic tangent model. *J. Mol. Liq* 231 (2017): 523-527.

17. Plazinski W, Rudzinski W, Kinetics of Adsorption at Solid/Solution Interfaces Controlled by Intraparticle Diffusion: A Theoretical Analysis. *J Phys Chem C* 113 (2009): 12495-12501.
18. Rudzinski W, Plazinski W. On the applicability of the pseudo-second order equation to represent the kinetics of adsorption at solid/solution interfaces: a theoretical analysis based on the statistical rate theory. *Adsorption* 15 (2009): 181-192.
19. Isiuku BO, Anyanwu DC, Nwosu CN, et al. Batch adsorptive removal of metanil yellow from simulated dye wastewater using nitric acid modified-H₃PO₄-activated carbon (NAMPAAC) from Gmelina arborea bark: equilibrium and isotherm studies. *Desalin Water Treat* 70 (2017): 380-386.
20. Mahmoodi NM, Salehi R, Arami M. Binary system dye removal from colored textile wastewater using activated carbon: Kinetic and isotherm studies. *Desalination* 272 (2011): 187-195.
21. Malik PK, Dye removal from wastewater using activated carbon developed from sawdust: adsorption equilibrium and kinetics. *J. Hazard. Mater* 113 (2004): 81-88.
22. Wang CT, Chou WL, Chung MH, et al. COD removal from real dyeing wastewater by electro-Fenton technology using an activated carbon fiber cathode. *Desalination* 253 (2010): 129-134.
23. Reddy DD, Ghosh RK, Bindu JP, et al. Removal of Methylene Blue from Aqueous System Using Tobacco Stems Biomass: Kinetics, Mechanism and Single-Stage Adsorber Design. *Environ Prog Sustain* 2017, 36 (4): 1005-1012.
24. Guo YF, Deng J, Zhu JY, et al. Removal of mercury(II) and methylene blue from a wastewater environment with magnetic graphene oxide: adsorption kinetics, isotherms and mechanism. *Rsc Adv* 6 (2016): 82523-82536.
25. Sen, T. K, Afroze, S, Ang, H. M, Equilibrium, Kinetics and Mechanism of Removal of Methylene Blue from Aqueous Solution by Adsorption onto Pine Cone Biomass of *Pinus radiata*. *Water Air Soil Poll* 2011, 218 (1-4): 499-515.
26. Kannan N, Sundaram MM. Kinetics and mechanism of removal of methylene blue by adsorption on various carbons - a comparative study. *Dyes Pigments* 51 (2001): 25-40.
27. Nasernejad B, Zadeh TE, Pour BB, et al. Comparison for biosorption modeling of heavy metals (Cr(III): Cu(II): Zn(II)) adsorption from wastewater by carrot residues. *Process Biochem* 40 (2005): 1319-1322.
28. Rengaraj S, Moon SH. Kinetics of adsorption of Co(II) removal from water and wastewater by ion exchange resins. *Water Res* 36 (2002): 1783-1793.
29. Raghunath S, Anand K, Gengan RM, et al. Sorption isotherms, kinetic and optimization process of amino acid proline based polymer nanocomposite for the removal of selected textile dyes from industrial wastewater. *J Photoch Photobio B* 165 (2016): 189-201.
30. Khorramfar S, Mahmoodi NM, Arami M, et al. Equilibrium and kinetic studies of the cationic dye removal capability of a novel biosorbent *Tamarindus indica* from textile wastewater. *Color Technol* 126 (2010): 261-268.
31. Al-Ghouti MA, Khraisheh MA, Allen SJ, et al. The removal of dyes from textile wastewater: a study of the physical characteristics and adsorption mechanisms of diatomaceous earth. *J. Environ. Manage* 69 (2003): 229-238.

32. Ho YS, McKay G. Pseudo-second order model for sorption processes. *Process Biochem* 34 (1999): 451-465.
33. Ho YS, McKay G. A comparison of chemisorption kinetic models applied to pollutant removal on various sorbents. *Process Saf. Environ. Prot* 76 (1998): 332-340.
34. Ho YS. Review of second-order models for adsorption systems. *J. Hazard. Mater* 136 (2006): 681-689.
35. Demirbas E, Kobya M, Senturk E, et al. Adsorption kinetics for the removal of chromium (VI) from aqueous solutions on the activated carbons prepared from agricultural wastes. *Water Sa* 30 (2004): 533-539.
36. Sun QY, Yang LZ. The adsorption of basic dyes from aqueous solution on modified peat-resin particle. *Water Res* 37 (2003): 1535-1544.
37. Hameed BH, Din ATM, Ahmad AL. Adsorption of methylene blue onto bamboo-based activated carbon: Kinetics and equilibrium studies. *J. Hazard. Mater* 141 (2007): 819-825.
38. Qiu H, Lv L, Pan BC, et al. Critical review in adsorption kinetic models. *J Zhejiang Univ-Sc A* 10 (2009): 716-724.
39. Putra EK, Pranowo R, Sunarso J, Indraswati, N, Ismadji, S, Performance of activated carbon and bentonite for adsorption of amoxicillin from wastewater: Mechanisms, isotherms and kinetics. *Water Res* 43 (2009): 2419-2430.
40. Carroll KM, Harkness MR, Bracco AA, et al. Application of a Permeant Polymer Diffusional Model to the Desorption of Polychlorinated-Biphenyls from Hudson River Sediments. *Environ. Sci. Technol* 28 (1994): 253-258.
41. Pignatello JJ, Ferrandino FJ, Huang LQ. Elution of Aged and Freshly Added Herbicides from a Soil. *Environ. Sci. Technol* 27 (1993): 1563-1571.
42. Steinberg SM, Pignatello JJ, Sawhney BL. Persistence of 1,2-Dibromoethane in Soils - Entrapment in Intraparticle Micropores. *Environ. Sci. Technol* 21 (1987): 1201-1208.
43. Juang RS, Shiau JY. Adsorption isotherms of phenols from water onto macroreticular resins. *J. Hazard. Mater* 70 (1999): 171-183.
44. Kanazaki T, Hirawa S, Harada M, et al. Coupled Acoustic-Gravity Field for Dynamic Evaluation of Ion Exchange with a Single Resin Bead. *Anal. Chem* 82 (2010): 4472-4478.
45. Kukkonen JVK, Landrum PF, Mitra S, et al. Sediment characteristics affecting desorption kinetics of select PAH and PCB congeners for seven laboratory spiked sediments. *Environ. Sci. Technol* 37 (2003): 4656-4663.
46. Naef R, Acree WE. Calculation of Five Thermodynamic Molecular Descriptors by Means of a General Computer Algorithm Based on the Group-Additivity Method: Standard Enthalpies of Vaporization, Sublimation and Solvation, and Entropy of Fusion of Ordinary Organic Molecules and Total Phase-Change Entropy of Liquid Crystals. *Molecules* 22 (2017).
47. Chen MS, Fan HF, Lin KC. Kinetic and thermodynamic investigation of rhodamine B adsorption at solid/solvent interfaces by use of evanescent-wave cavity ring-down spectroscopy. *Anal. Chem* 82 (2010): 868-877.

48. Wegner GJ, Wark AW, Lee HJ, et al. Real-time surface plasmon resonance imaging measurements for the multiplexed determination of protein adsorption/desorption kinetics and surface enzymatic reactions on peptide microarrays. *Anal. Chem* 76 (2004): 5677-5684.
49. Nkedikizza P, Brusseau ML, Rao PSC, et al. Nonequilibrium Sorption during Displacement of Hydrophobic Organic-Chemicals and Ca-45 through Soil Columns with Aqueous and Mixed-Solvents. *Environ. Sci. Technol* 23 (1989): 814-820.
50. Llewellyn PL, Maurin G, Devic T, et al. Prediction of the conditions for breathing of metal organic framework materials using a combination of X-ray powder diffraction, microcalorimetry, and molecular simulation. *JACS* 130 (2008): 12808-12814.
51. Bourrelly S, Llewellyn PL, Serre C, et al. Different adsorption behaviors of methane and carbon dioxide in the isotypic nanoporous metal terephthalates MIL-53 and MIL-47. *JACS* 127 (2005): 13519-13521.
52. Sapsford KE, Tyner KM, Dair BJ, Deschamps, J. R, Medintz, I. L, Analyzing Nanomaterial Bioconjugates: A Review of Current and Emerging Purification and Characterization Techniques. *Anal. Chem* 83 (2011): 4453-4488.
53. Pignatello JJ, Xing BS. Mechanisms of slow sorption of organic chemicals to natural particles. *Environ. Sci. Technol* 30 (1996): 1-11.
54. Xing BS, Pignatello JJ. Dual-mode sorption of low-polarity compounds in glassy poly(vinyl chloride) and soil organic matter. *Environ. Sci. Technol* 31 (1997): 792-799.
55. Salles F, Ghoufi A, Maurin G, et al. Molecular Dynamics Simulations of Breathing MOFs: Structural Transformations of MIL-53(Cr) upon Thermal Activation and CO₂ Adsorption. *Angew Chem Int Edit* 47 (2008): 8487-8491.
56. Farrell J, Grassian D, Jones M. Investigation of mechanisms contributing to slow desorption of hydrophobic organic compounds from mineral solids. *Environ. Sci. Technol* 33 (1999): 1237-1243.
57. Berens AR, Hopfenberg HB. Diffusion and Relaxation in Glassy Polymer Powders .2. Separation of Diffusion and Relaxation Parameters. *Polymer* 19 (1978): 489-496.
58. Berens AR. Diffusion and Relaxation in Glassy Polymer Powders .1. Fickian Diffusion of Vinyl-Chloride in Polyvinyl Chloride). *Polymer* 18 (1977): 697-704.
59. Connaughton DF, Stedinger JR, Lion LW, et al. Description of Time-Varying Desorption-Kinetics - Release of Naphthalene from Contaminated Soils. *Environ. Sci. Technol* 27 (1993): 2397-2403.
60. Grathwohl P, Reinhard M. Desorption of Trichloroethylene in Aquifer Material - Rate Limitation at the Grain Scale. *Environ. Sci. Technol* 27 (1993): 2360-2366.
61. Harmon TC, Roberts PV. Comparison of Intraparticle Sorption and Desorption Rates for a Halogenated Alkene in a Sandy Aquifer Material. *Environ. Sci. Technol* 28 (1994): 1650-1660.
62. Berens AR, Huvard GS. Particle-Size Distribution of Polymer Powders by Analysis of Sorption Kinetics. *J. Dispersion Sci. Technol* 2 (1981): 359-378.
63. Lou S, Chen ZB, Liu YF, et al. Synthesis of Functional Adsorption Resin and Its Adsorption Properties in Purification of Flavonoids from *Hippophae rhamnoides* L. Leaves. *Industrial & Engineering Chemistry Research* 51 (2012): 2682-2696.

64. Lou S, Di D. Synthesis of resins with ionic liquids for purification of flavonoids from *Hippophae rhamnoides* L. leaves. *J. Agric. Food. Chem* 60 (2012): 6546-6558.
65. Lou S, Chen Z, Liu Y, et al. New way to analyze the adsorption behavior of flavonoids on macroporous adsorption resins functionalized with chloromethyl and amino groups. *Langmuir* 27 (2011): 9314-9326.
66. Dallabetta RA. Carbon-Monoxide Adsorption on Supported Ruthenium. *J. Phys. Chem* 79 (1975): 2519-2525.
67. Lasakova M, Jandera P. Molecularly imprinted polymers and their application in solid phase extraction. *J. Sep. Sci* 32 (2009): 799-812.
68. Lou S, Liu YF, Bai QQ, et al. Adsorption Mechanism of Macroporous Adsorption Resins. *Prog Chem* 24 (2012): 1427-1436.
69. Wegner GJ, Wark AW, Lee HJ, et al. Real-time surface plasmon resonance imaging measurements for the multiplexed determination of protein adsorption/desorption kinetics and surface enzymatic reactions on peptide microarrays. *Anal. Chem* 76 (2004): 5677-5684.



This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC-BY) license 4.0